

FORMATIVE ASSESSMENT-3

Name:..... Section:..... Roll No:..... Max.Marks:20

I. Answer the following questions. Each carries four marks. 2 x 4 = 8 M

- 1) Rainbow is an example for continuous spectrum. Explain.
- 2) What do you understand about the periodic properties of elements. Complete the following table.

Property of elements	Trend in group	Trend in Period
	(Increased/Decreased)	(Increased/Decreased)
Atomic radius		
Ionisation energy		
Electro negativity		
Metallic nature		

II. Answer the following questions briefly. Each carries two marks. 2 x 2 = 4 M

- 3) Write the differences between Orbit and orbital.
- 4) Write the following properties of element having atomic number 16.
  - (i) Electron configuration
  - (ii) Valence electrons
  - (iii) Valency
  - (iv) Is it a metal or non metal ?

III. Answer the following in one or two sentences. Each carries one marks. 2 x 1 = 2 M

- 5) Write Planck's equation.
- 6) Give one example for Dobernier's triads.

IV. Choose the correct choice and write down in the given brackets. 6 x 1 = 6 M

- 7) If  $l = 0$ , the sub shell is ..... [ ]
  - A. s
  - B. p
  - C. d
  - D. f
- 8) The colour of flame produced when  $CuCl_2$  is burnt [ ]
  - A. Red
  - B. Yellow
  - C. Brown
  - D. Blue
- 9) Correct electron configuration of Nitrogen [ ]
  - A.  $1s^2 2s^2 2p^4$
  - B.  $1s^2 2s^0 2p^5$
  - C.  $1s^2 2s^2 2p^3$
  - D.  $1s^2 2s^5$
- 10) "If I hold it in my hand, it will melt." – What is it ? [ ]
  - A. Chlorine
  - B. Gallium
  - C. Chromium
  - D. Mercury
- 11) Electro negativity = ..... [ ]
  - A.  $\frac{\text{Ionisation energy} + \text{Electron affinity}}{4}$
  - B.  $\frac{\text{Ionisation energy} - \text{Electron affinity}}{4}$
  - C.  $\frac{\text{Ionisation energy} + \text{Electron affinity}}{2}$
  - D.  $\frac{\text{Ionisation energy} - \text{Electron affinity}}{2}$
- 12) Identify the wrong statement. [ ]
  - A. Helium is an inert gas
  - B. Oxygen is a chalcogen
  - C. Sodium is a alkali metal
  - D. Cadmium is the liquid metal

